

Chemistry / Environmental Studies L111
Environmental Concerns and Chemical Solutions
Professor Dransfield
Homework 3: Due February 22, 2007

Chapter 2 – Protecting the Ozone Layer

Solve this problem:

We have learned that diatomic oxygen absorbs light with wavelengths less than 242 nm, and that ozone absorbs light with wavelengths less than 320 nm. What is the energy required to break the bonds in each of these molecules, in Joules? Can you explain these results using Lewis structures?

And the following questions from the text:

Fifth Edition Question Numbers:

Concentrating on Concepts - #34, 39, 43

Exploring Extensions - #53

Fourth Edition Question Numbers:

Concentrating on Concepts - #30, 35, 37

Exploring Extensions - #43