

Name: _____

CHM 115 Quiz #4

1) (5 pts) Write the oxidation number for the underlined element in:

a) H_{2(g)} _____ 0 _____

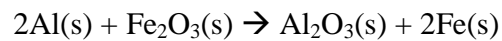
b) KClO_{2(aq)} _____ +3 _____

c) H₂O_(l) _____ +1 _____

d) Fe⁺³ _____ +3 _____

e) H₂SO₄ _____ -2 _____

2) (2 pts) For the following reaction, identify which element is reduced and which is oxidized:



Reduced: _____ Fe _____

Oxidized: _____ Al _____

3) (3 pts) How much 1M HCl stock solution is required to make 100 mL of 0.2M HCl?

$$C_i V_i = C_f V_f$$

$$V_i = (C_f V_f) / C_i$$

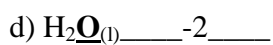
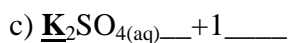
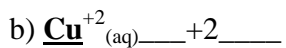
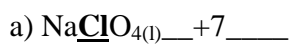
$$V_i = (0.2 \text{ M HCl})(100 \text{ mL}) / (1 \text{ M HCl})$$

$$V_i = 20 \text{ mL HCl}$$

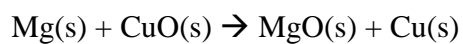
Name: _____

CHM 115 Quiz #4

1) (5 pts) Write the oxidation number for the **underlined** element in:



2) (2 pts) For the following reaction, identify which element is reduced and which is oxidized:



Reduced: _____ Cu _____

Oxidized: _____ Mg _____

3) (3 pts) How much 2M HCl stock solution is required to make 50 mL of 0.2M HCl?

$$C_i V_i = C_f V_f$$

$$V_i = (C_f V_f) / C_i$$

$$V_i = (0.2 \text{ M HCl})(50 \text{ mL}) / (2 \text{ M HCl})$$

$$V_i = 5 \text{ mL HCl}$$