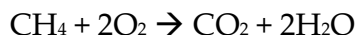


Quiz #3

Name: \_\_\_\_\_

1) If you have 4 million molecules of CH<sub>4</sub> and 2 million molecules of O<sub>2</sub>, and they react as show:



Which reactant is limiting? \_\_\_O<sub>2</sub>\_\_\_\_\_

And how many molecules of CO<sub>2</sub> will be produced? \_\_\_\_\_1 Million\_\_\_\_\_

2) In a solution of water containing NaCl, what is the solvent and what is the solute:

solvent:\_\_\_WATER\_\_\_\_\_

soluted:\_\_\_\_\_NaCl\_\_\_\_\_

3) Three aqueous solutions are made using three different salts: Na<sub>3</sub>PO<sub>4</sub>, K<sub>2</sub>SO<sub>4</sub>, and NaCl.

In terms of electrical conductivity, the solution made with \_\_\_Na<sub>3</sub>PO<sub>4</sub> (most ions)\_\_\_\_\_ will be the most conductive and the solution containing \_\_\_NaCl (least ions)\_\_\_\_\_ will be the least conductive.

4) Circle the correct label:

ELECTROLYTE / non-electrolyte - KBr (ionic)

ELECTROLYTE / non-electrolyte – Ca(NO<sub>3</sub>)<sub>2</sub> (ionic)

5) Circle the word that best fits the sentence:

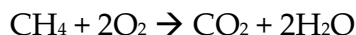
Compounds of NH<sub>4</sub><sup>+</sup> and NO<sub>3</sub><sup>-</sup> are always SOLUBLE/insoluble

Ag<sup>+</sup>, Hg<sub>2</sub><sup>+2</sup>, and Pb<sup>+2</sup> are often soluble / INSOLUBLE in comparison to most other metal salts.

Quiz #3

Name: \_\_\_\_\_

1) If you have 3 million molecules of CH<sub>4</sub> and 2 million molecules of O<sub>2</sub>, and they react as show:



Which reactant is limiting? \_\_\_\_O<sub>2</sub>\_\_\_\_\_

And how many molecules of CO<sub>2</sub> will be produced? \_\_\_\_1 Million\_\_\_\_\_

2) In a solution of water containing KBr, what is the solvent and what is the solute:

solute: \_\_\_\_KBr\_\_\_\_\_

solvent: \_\_\_\_Water\_\_\_\_\_

3) Three aqueous solutions are made using three different salts: Na<sub>2</sub>SO<sub>4</sub>, K<sub>3</sub>PO<sub>4</sub>, and KCl.

In terms of electrical conductivity, the solution made with \_\_\_\_K<sub>3</sub>PO<sub>4</sub> (most ions)\_\_\_\_\_ will be the most conductive and the solution containing \_\_\_\_KCl (least ions)\_\_\_\_\_ will be the least conductive.

4) Circle the correct label:

ELECTROLYTE / non-electrolyte - K<sub>2</sub>SO<sub>4</sub>

ELECTROLYTE / non-electrolyte - HCl

5) Circle the word that best fits the sentence:

Ag<sup>+</sup>, Hg<sub>2</sub><sup>+2</sup>, and Pb<sup>+2</sup> are often soluble / INSOLUBLE in comparison to most other metal salts.

Compounds of NH<sub>4</sub><sup>+</sup> and NO<sub>3</sub><sup>-</sup> are always SOLUBLE / insoluble.