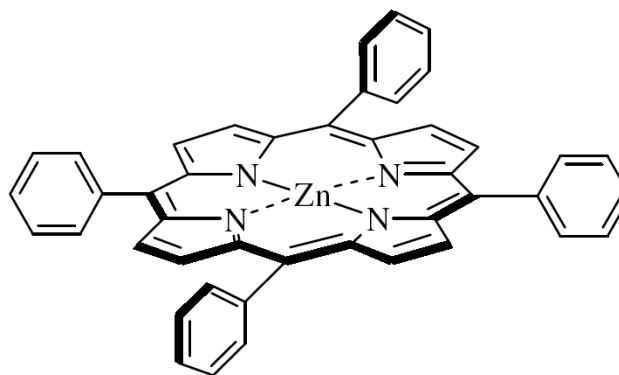
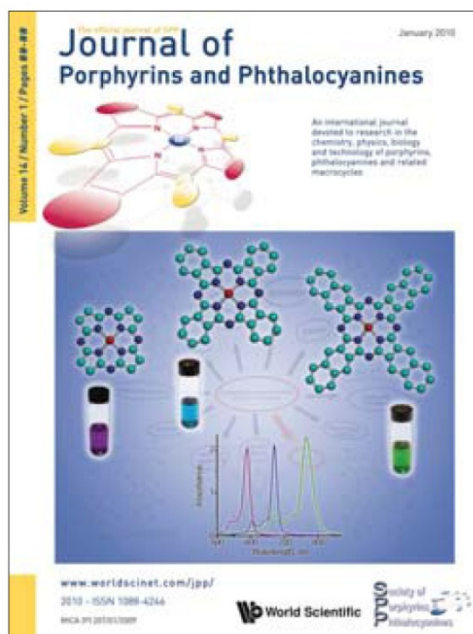
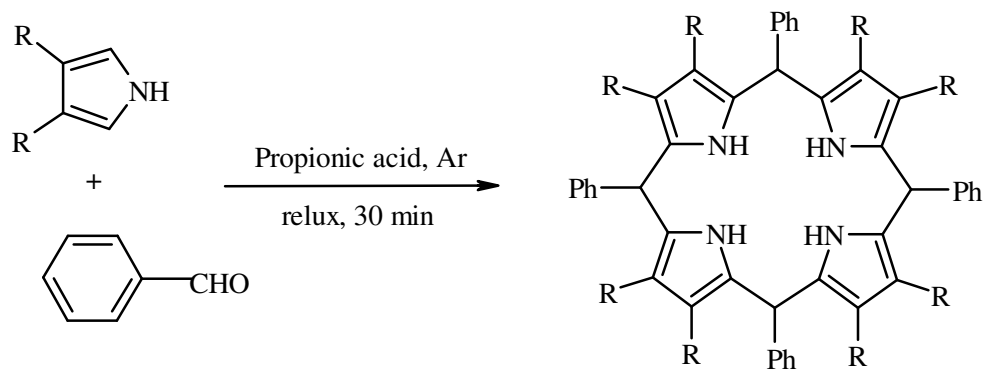


Microscale Synthesis of Porphyrin Complexes



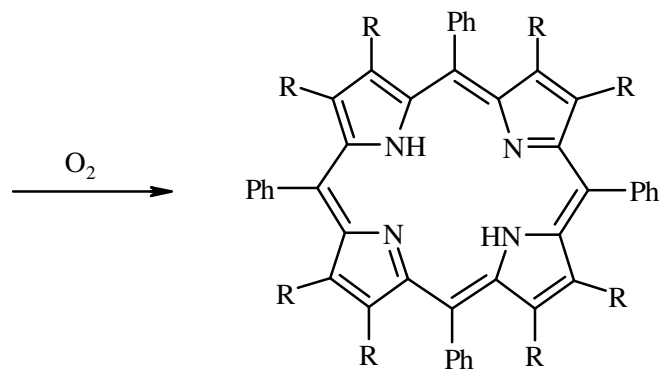
Synthetic Methods – Adler method



R = H, Me

R = H - meso-Tetraphenylporphyrinogen

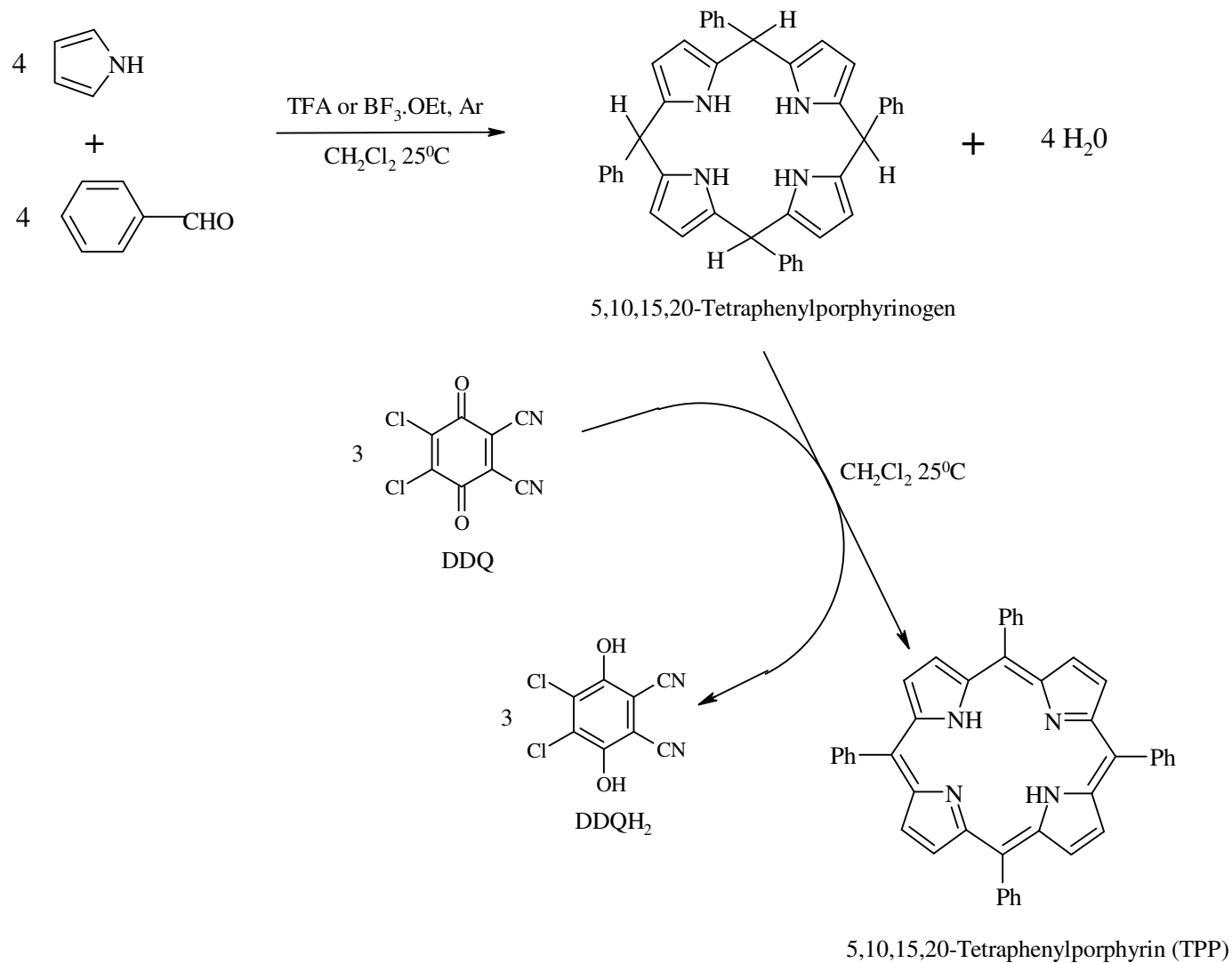
R = Me - β -octamethyl-meso-tetraphenylporphyrinogen



R = H - meso-Tetraphenylporphyrin

R = Me - β -octamethyl-meso-tetraphenylporphyrin

Synthetic Methods – Lindsey method

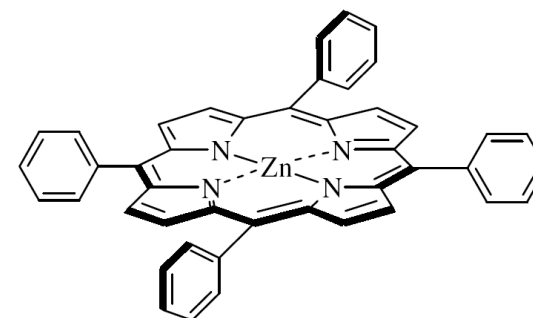


The Porphyrin Chromophore

Porphyrins are particularly attractive for many applications due to their

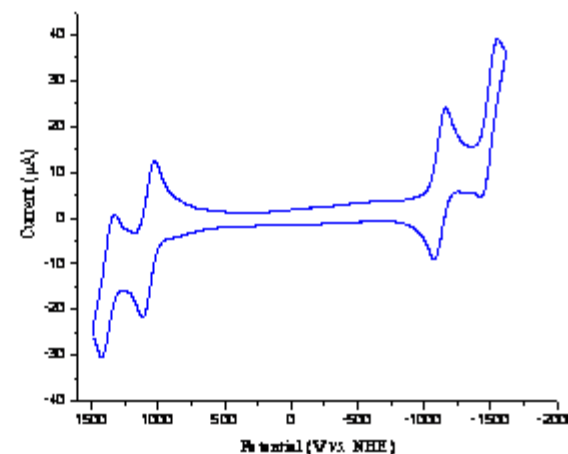
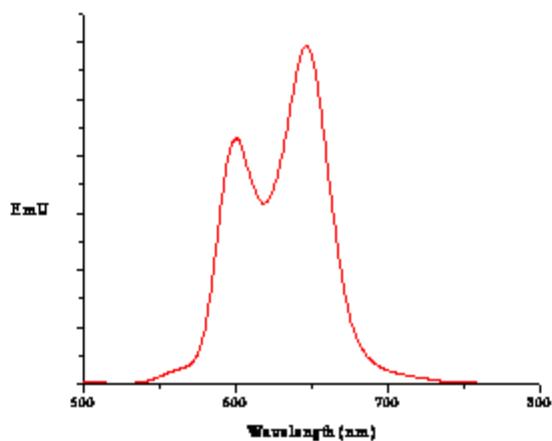
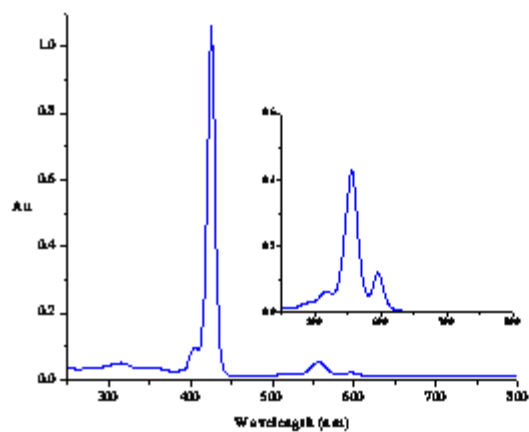
- rigid-planar geometry
- redox stability
- intense electronic absorption
- visible emission
- small HOMO-LUMO energy gap
- tunable optical and redox properties

Zinc-5,10,15,20-tetraphenylporphyrin



UV-Vis absorption		Fluorescence	
Soret	Q(1,0)	Q(0,0)	Q(0,0)*, Q(1,0)*
λ_{\max} , nm	λ_{\max} , nm	λ_{\max} , nm	λ_{\max} , nm
($\epsilon \times 10^5$, M ⁻¹ cm ⁻¹)	($\epsilon \times 10^5$, M ⁻¹ cm ⁻¹)	($\epsilon \times 10^5$, M ⁻¹ cm ⁻¹)	(Φ)
422 (2.55)	548 (9.72)	588 (1.90)	596, 647 (0.033)

Oxidation		Reduction		E_{0-0}	$E_{1/2}(P^+/P^*)$
(V)		(V)		(eV)	(eV)
1 st	2 nd	1 st	2 nd		
1.10	1.46	-1.14	-1.49	2.07	-0.97

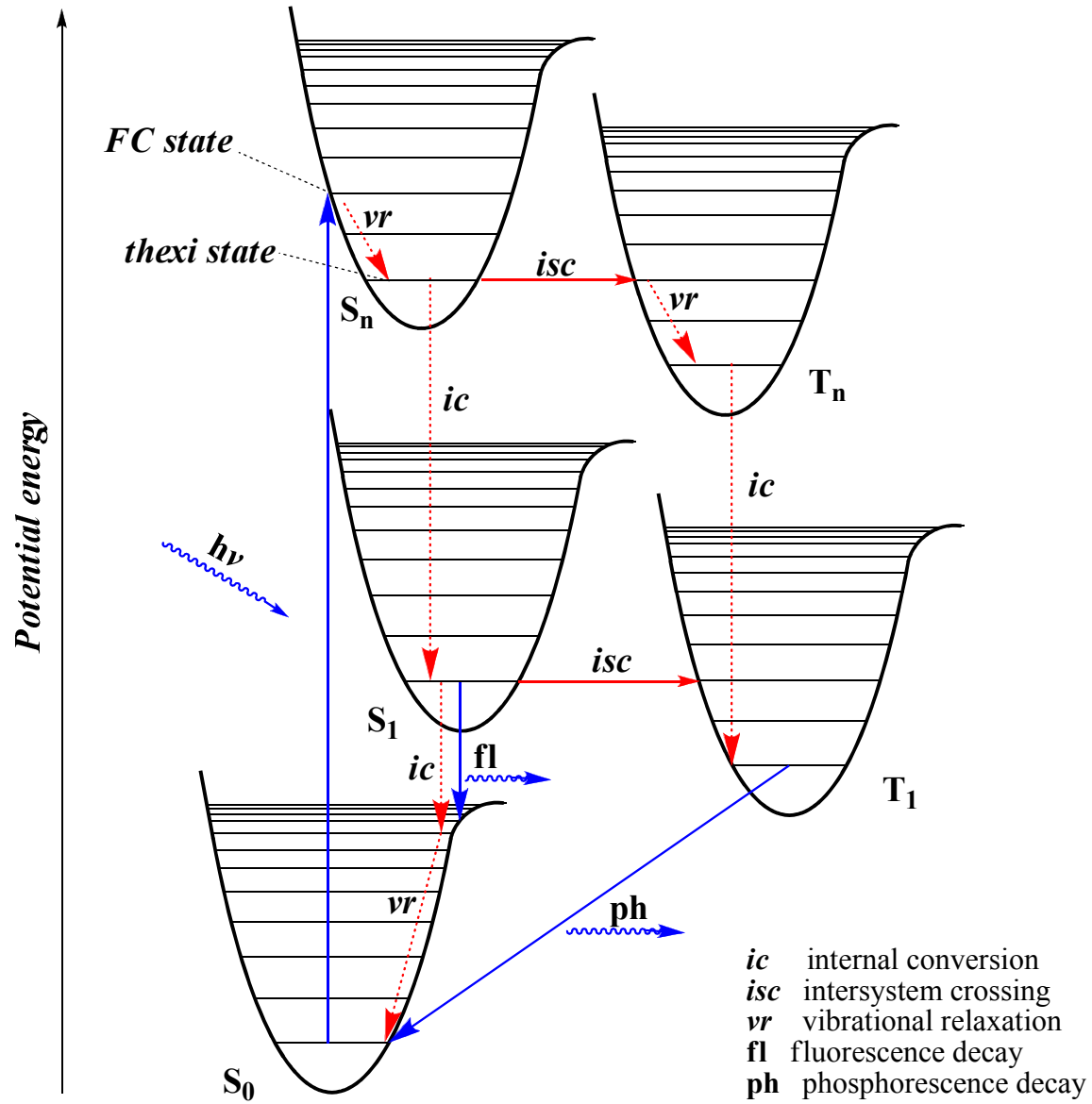


UV-Vis absorption, Fluorescence Emission and Cyclic Voltammetry of ZnTPP

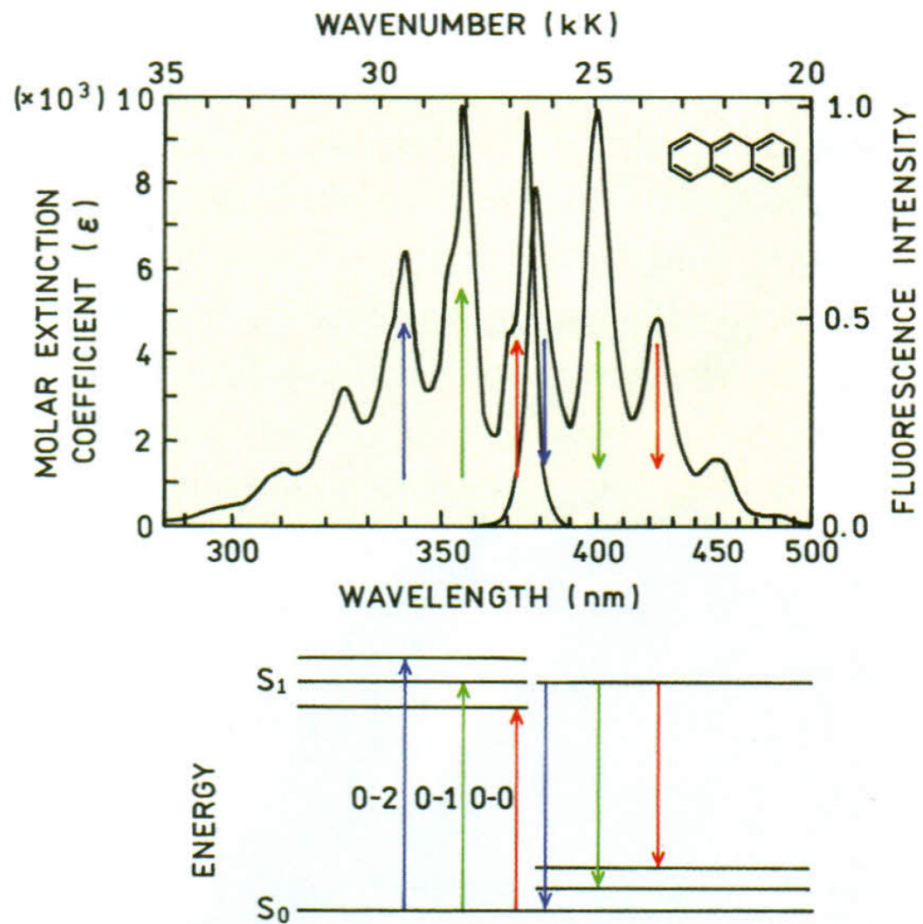
Introduction to Photophysics

- Upon irradiation, the a chromophore is excited from its (typically) singlet ground state (S_0) to a high energy Frank-Condon state.
- The Frank-Condon state undergoes rapid internal conversion and vibrational relaxation (often referred to as radiationless decay) to the first singlet excited state (S_1) at lower energy [sometimes referred to as a **thermally-equilibrated excited (thexi) state**].
- A number of energy loss mechanisms are possible from the S_1 state of the PS:
 - **Fluorescence emission (*nanosecond timescale*).**
 - **Internal conversion (*ic*) and vibrational relaxation (*radiationless decay*).**
 - **Intersystem crossing (*isc*) to the triplet excited state T_1 .**
- From the T_1 state *phosphorescence emission* (micro- to millisecond timescale) occurs, or in the presence of O_2 reaction may occur via type 1 or type 2 mechanisms producing reactive oxygen species (re: photodynamic therapy).

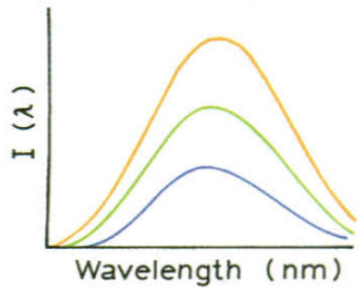
Jablonski diagram



Fluorescence Spectroscopy (e.g. anthracene)



- Mirror-image rule and Franck-Condon factors.
- The numbers 0, 1, and 2 refer to vibrational energy levels.
- ***As phosphorescence is a spin forbidden $T_1 - S_0$ transition the mirror rule does not apply.***



$$\Phi_f = \text{no. of photons emitted} / \text{no. of photons absorbed}$$

- Fluorescence quantum yields (Φ_f) are calculated using *steady state* methods by **actinometry**:

Typically, in de-aerated, optically dilute solutions fluorescence spectra of both the actinometer (ref) and sample (s) are recorded following identical monochromatic excitation of uniform intensity

$$\Phi_f = (A_{\text{ref}}/A_s)(I_s/I_{\text{ref}})(\eta_s/\eta_{\text{ref}})^2\Phi_{\text{ref}}$$

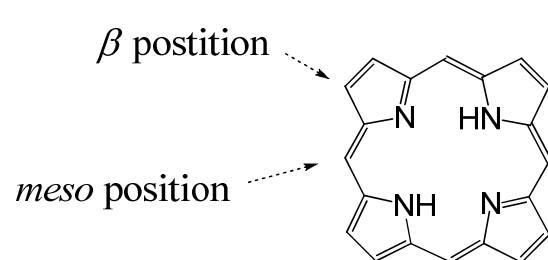
where A is the absorbance at the excitation wavelength, I is the integrated emission area and η is the solvent refractive index.

- If an actinometer is used which allows use of the same solvent this equation reduces to

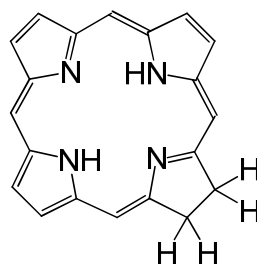
$$\Phi_f = (A_{\text{ref}}/A_s)(I_s/I_{\text{ref}})\Phi_{\text{ref}}$$

The Porphyrinoid Family

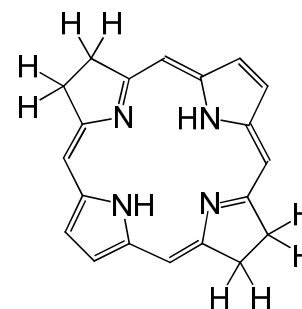
- The most extensively studied photosensitizers are *tetrapyrrole* chromophores, which includes porphyrins, chlorins, bacteriochlorins, benzoporphyrins, and phthalocyanines.



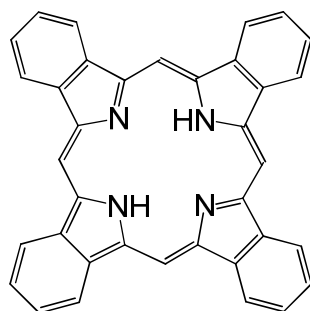
porphine



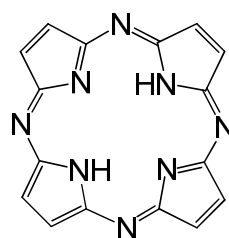
chlorin



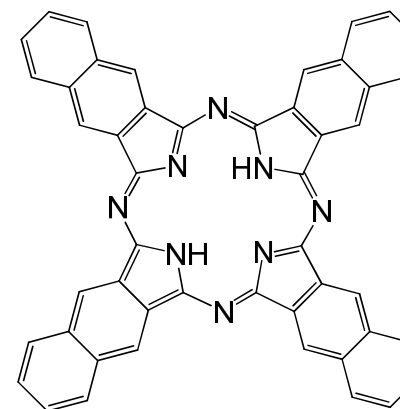
bacteriochlorin



benzoporphyrin

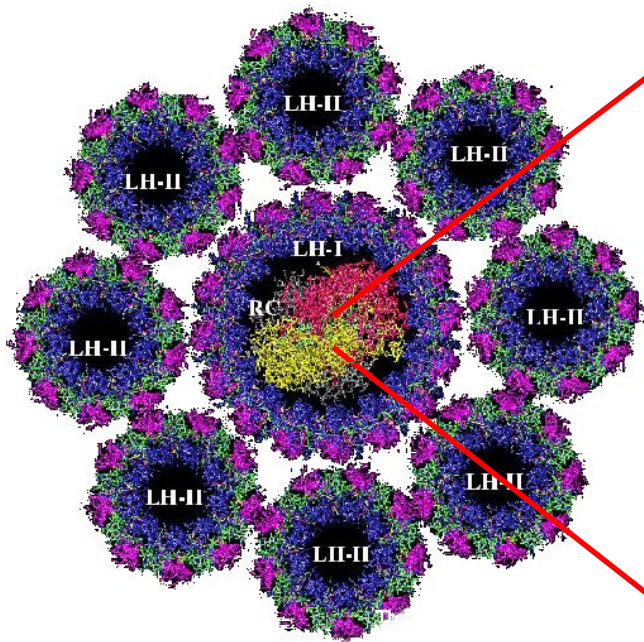


phthalocyanine

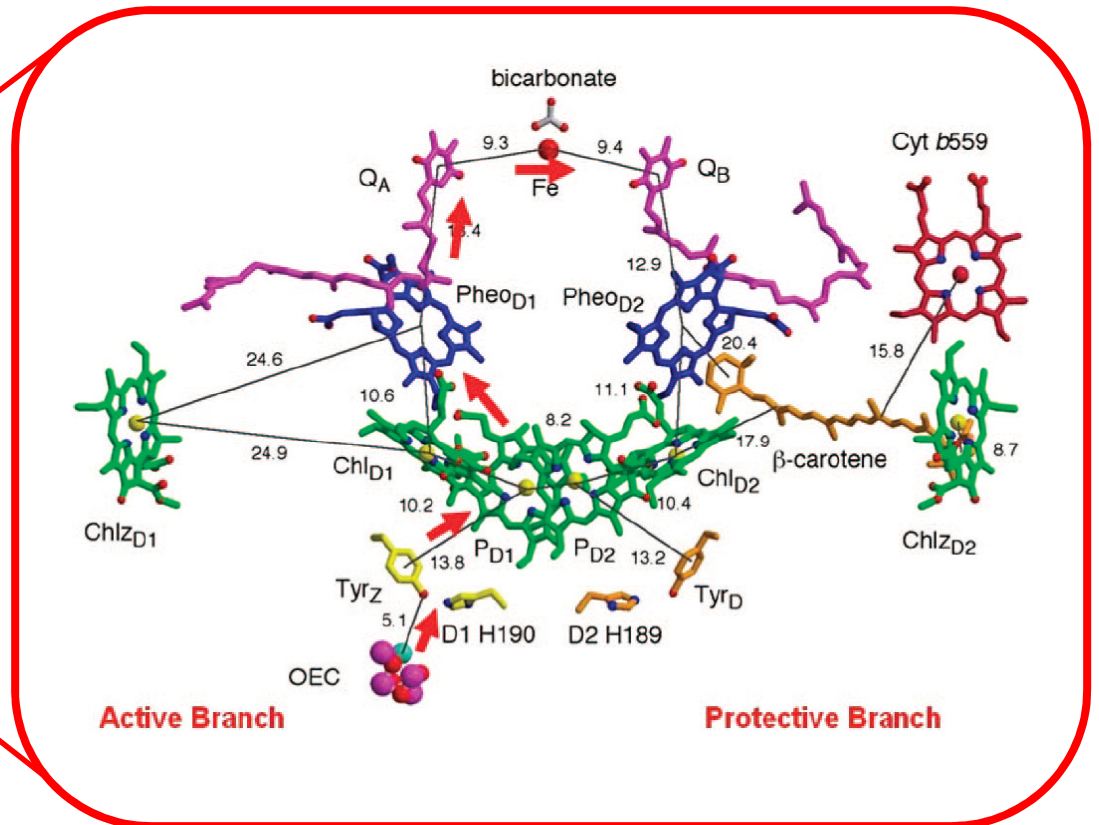


naphthalocyanine

Photosynthesis



Light-Harvesting
(energy transfer)



Reaction Center
(charge-separation via e^- transfer)

Porphyrin Catalysts for CO₂ Reduction

- The metalloporphyrins and related metallo-macrocycles studied for CO₂ reduction reactivity include metalloporphyrins (MP), metallocorrins (MN), metallophthalocyanines (MPc), and metallocorroles (MC), Figure 4, where M = Fe or Co.
- The active catalytic states as identified by cyclic voltametry have the metal in the formal oxidation state of zero for porphyrins [M⁰P]²⁻ and corrins [M⁰N]²⁻, +1 for corroles [M⁺C]²⁻, and +1 with a reduced phthalocyanine ring [M⁺Pc^{•-}]²⁻.

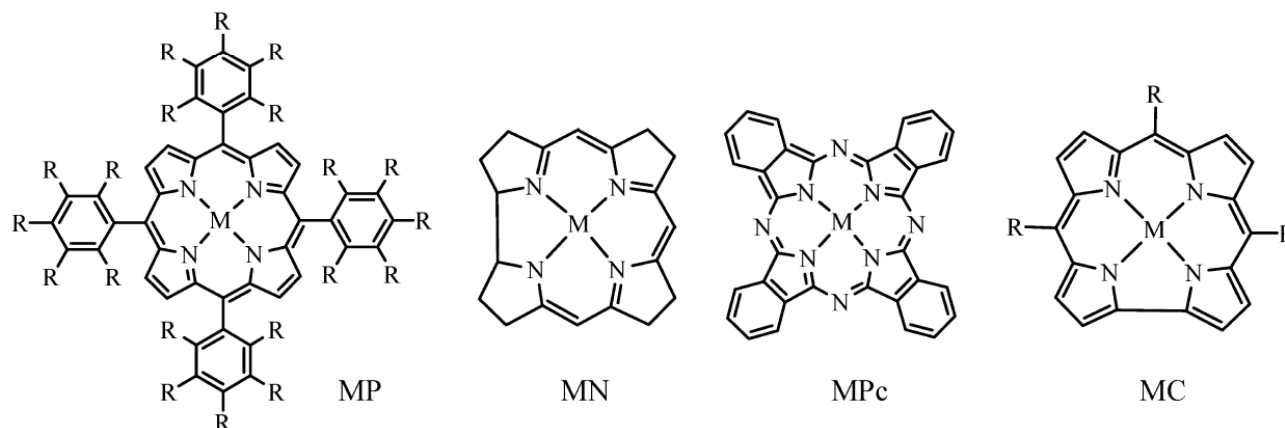
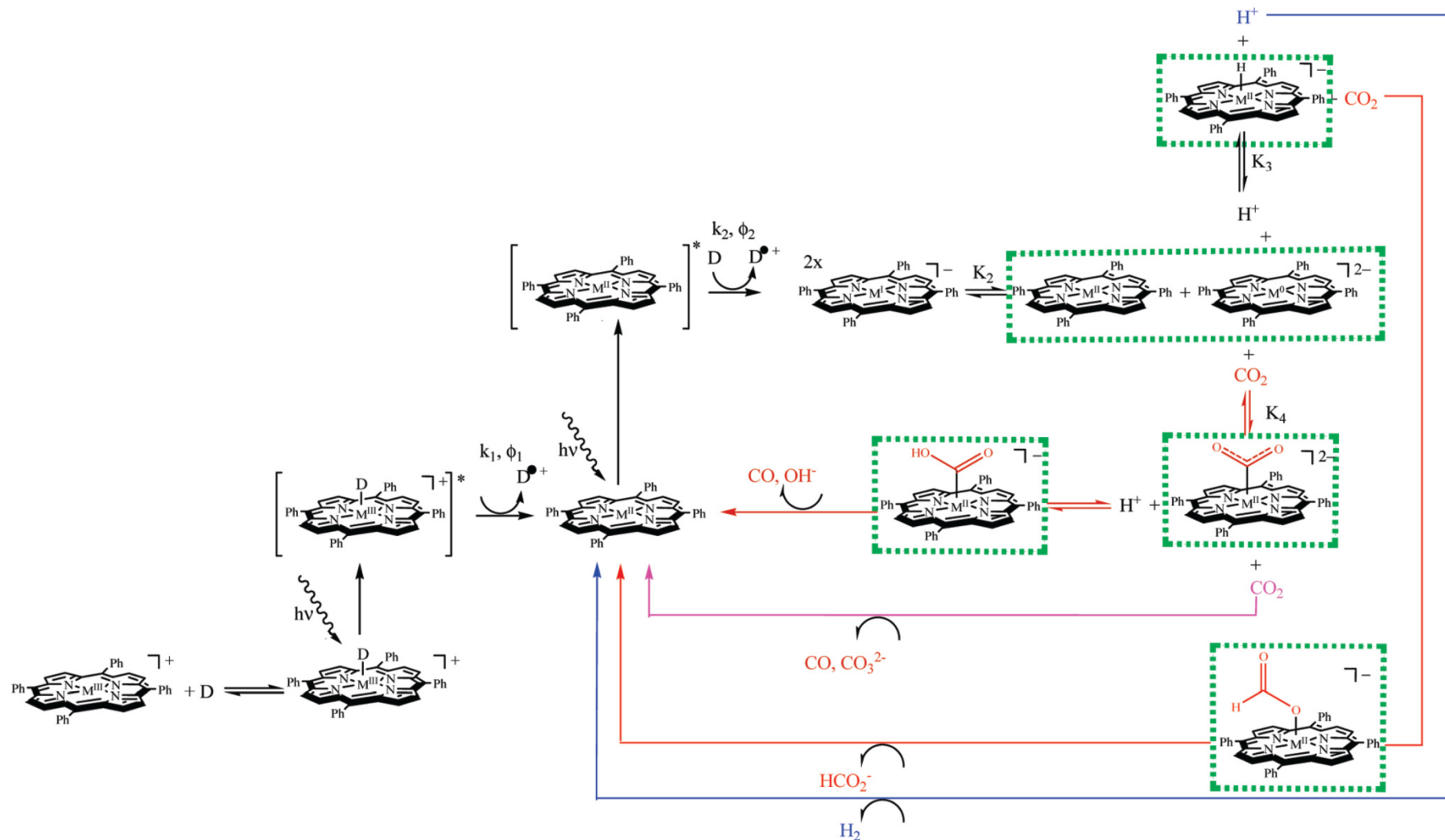


FIGURE 4. Metal porphyrin derivatives investigated for CO₂ reduction: (left to right) metalloporphyrin (MP), metallocorrin (MN), metallophthalocyanine (MPc), and metallocorrole (MC, where R = C₆F₅ or 2,6-C₆H₃Cl₂).

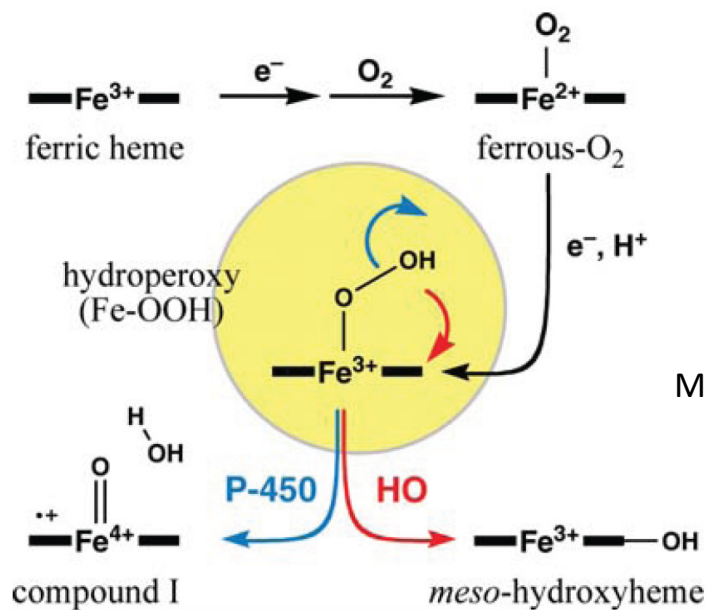
Mechanism for CO₂ Reduction

SCHEME 2. Proposed Mechanistic Steps in the Reduction of CO₂ by Metal Porphyrin Derivatives (M = Fe or Co) via a Type II Mechanism^a



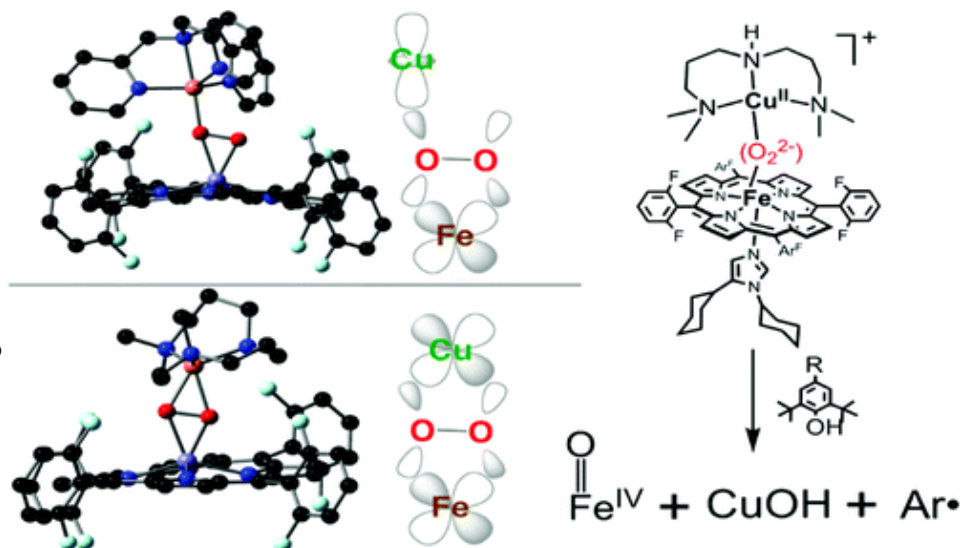
^a Hydrogen production (blue), formate production (red), CO formation (brown and pink), and putative intermediates (green); as a representative compound, metalloporphyrin is illustrated.

Porphyrin Catalysts for O₂ Activation



Matsui et. al *Inorg. Chem.*, 2010, 49 (8), pp 3602–3609

Figure 2. Comparison of O₂ activation by HO and cytochrome P450.



Halime et. al *Inorg. Chem.*, 2010, 49 (8), pp 3629–3645

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doi:10.1038/nrc1894

Photodynamic therapy and anti-tumour immunity

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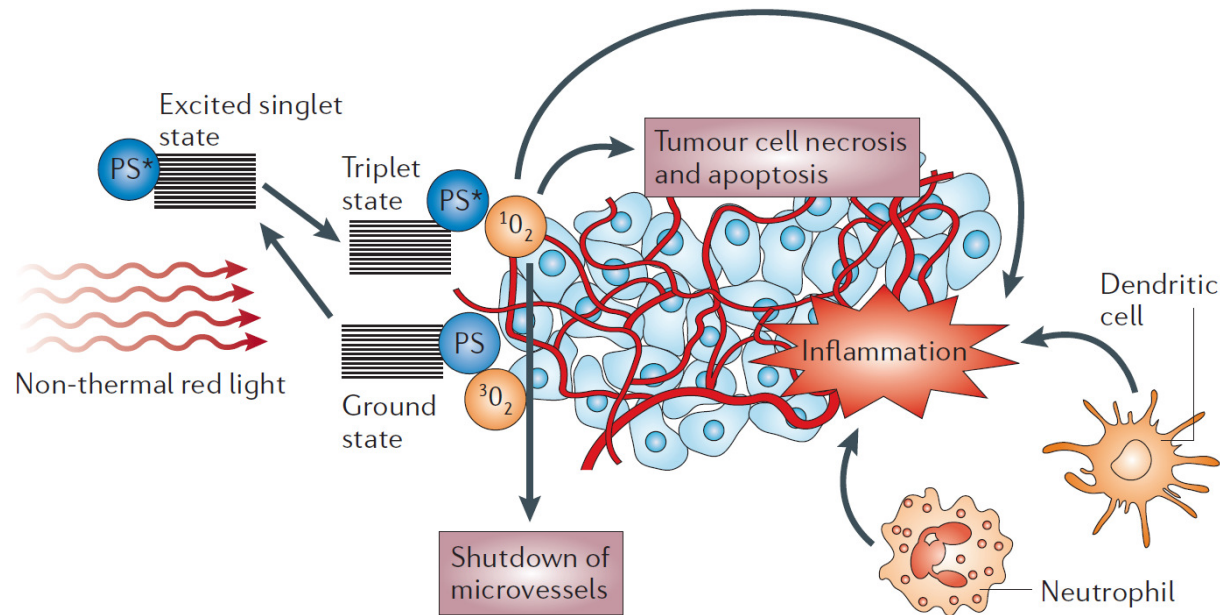
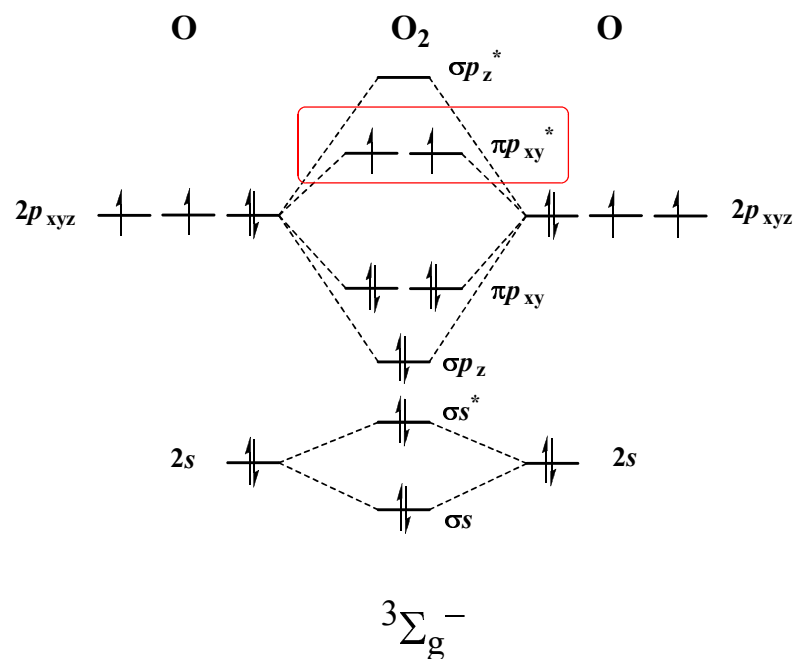
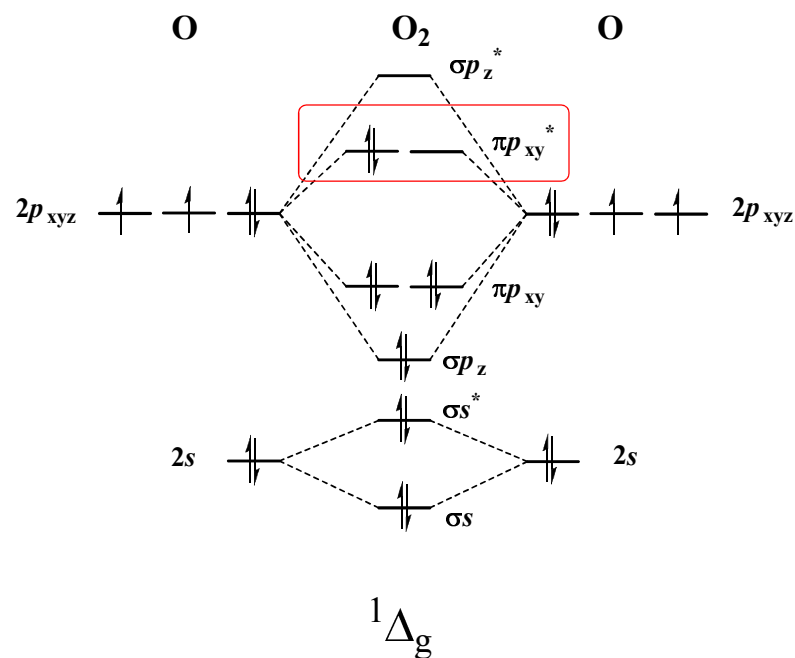
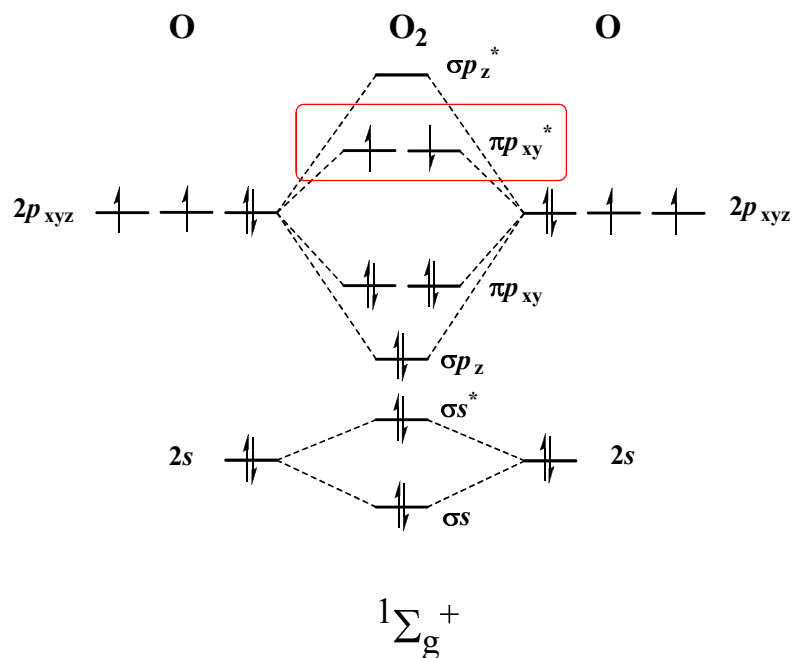


Figure 1 | **The mechanism of action on tumours in photodynamic therapy.**

The photosensitizer (PS) absorbs light and an electron moves to the first short-lived excited singlet state. This is followed by intersystem crossing, in which the excited electron changes its spin and produces a longer-lived triplet state. The PS triplet transfers energy to ground-state triplet oxygen, which produces reactive singlet oxygen (1O_2). 1O_2 can directly kill tumour cells by the induction of necrosis and/or apoptosis, can cause destruction of tumour vasculature and produces an acute inflammatory response that attracts leukocytes such as dendritic cells and neutrophils.



Molecular orbital diagrams for the three electronic configurations of O_2 .

The triplet ground state $3\Sigma_g^-$

The singlet oxygen $a^1\Delta_g$ excited state

The singlet oxygen $b^1\Sigma_g^+$ excited state.

Note that the states only differ in the spin and the occupancy of two degenerate anti-bonding π_g^* orbitals.

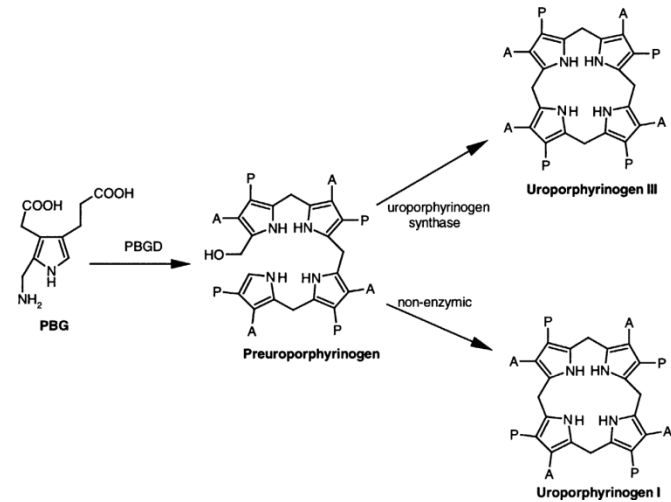
Friedrich Meyer-Betz, a German physician, injected himself with porphyrin in 1912 to test whether it made humans sensitive to light. The first photograph shows him four days later, after he took a walk on a sunny day. Most of the swelling subsided by the sixth day, second photograph, but he remained light-sensitive for several months.



courtesy A. M. McDonagh, UCSF (2)

Porphyria

- Overproduction and accumulation of the porphyrins (or their chemical precursors).
- Decreased production of heme leads to increased production of precursors, PBG being one of the first substances in the porphyrin synthesis pathway.
- The principal problem of porphyria is the accumulation of porphyrins, which are toxic to tissue in high concentrations.
- The chemical properties of these intermediates determine the location of accumulation, whether they induce photosensitivity.
- Porphyria has been suggested as an explanation for the **origin of vampire and werewolf legends**, based upon certain perceived similarities between the condition and the folklore.



- L. Illis' 1963 paper, "On Porphyria and the Aetiology of Werwolves", was published in *Proceedings of the Royal Society of Medicine*. In 1985, biochemist David Dolphin's paper for the American Association for the Advancement of Science, "Porphyria, Vampires, and Werewolves: The Aetiology of European Metamorphosis Legends", gained widespread media coverage, thus popularizing the connection.
- The theory has since faced heavy criticism, especially for the stigma it has placed on its sufferers. Norine Dresser's *American Vampires: Fans, Victims, Practitioners* (1989) treats the matter with more depth.