## Strong and Weak Acids and Bases

1. The following common acids are strong: HCl, HBr, HI, $HNO_3$, $HClO_4$, $H_2SO_4$

   The following are some less common acids that are also strong: $HClO_3$, $HBrO_3$, $HIO_3$, $H_2SeO_4$

   Assume all other acids are weak unless told otherwise.

   Some weak acids: HF, $HNO_2$, $HClO_2$, $[H_2SO_3] = SO_2 + H_2O$, $HC_2H_3O_2 = HOAc$

2. All ionic hydroxides are strong bases, regardless of solubility. Bases that do not contain $OH^-$ are weak.

   Some strong bases: LiOH, NaOH, KOH, RbOH, CsOH, $Mg(OH)_2$, $Ca(OH)_2$, $Sr(OH)_2$, $Ba(OH)_2$

   Some weak bases: $NH_3$, $CH_3NH_2$, $C_2H_5NH_2$

   Not a base: $CH_3OH$, $C_2H_5OH$