

Rules for Writing Ionic and Net Ionic Equations

1. Write all molecular compounds in molecular form.
Examples: $C_{11}H_{22}O_{11}(aq)$ $H_2O(l)$ $CO_2(g)$
2. Write all weak electrolytes in molecular form.
Examples: $CH_3CO_2H(aq)$ **Not** $H^+(aq) + CH_3CO_2^-(aq)$
 $NH_3(aq)$ **Not** $NH_4^+(aq) + OH^-(aq)$
3. Write soluble strong electrolytes as ions and insoluble strong electrolytes as combined solids (i.e., "molecular" form).
Examples: $NaCl(aq) = Na^+(aq) + Cl^-(aq)$
 $AgCl(s) = AgCl(s)$

Net Ionic Equations of Some Common Gas-Forming Reactions

