

Total valence electrons

$1 + 7 = 8$

$5 + 5 = 10$

$5 + 3(1) = 8$

$4 + 1 \times 4 = 8$

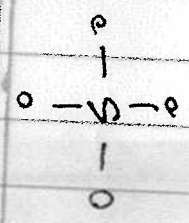
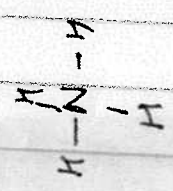
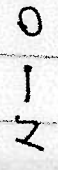
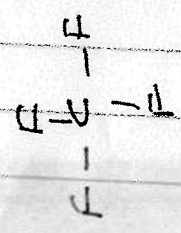
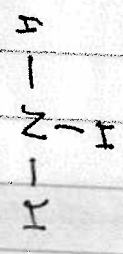
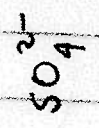
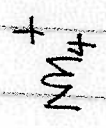
$4 + 4 \times 7 = 32$

$5 + 6 - 1 = 10$

$5 + 4 - 1 = 8$

$7 + 12 + 1 = 20e$

$6 + 6 \times 4 + 2 = 32e$



Remaining electrons

6

8

2

0

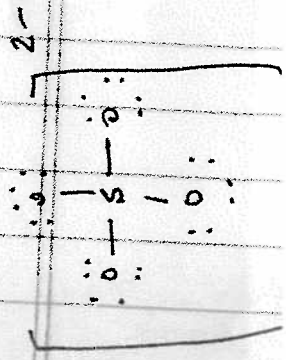
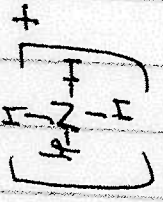
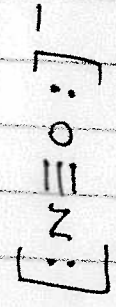
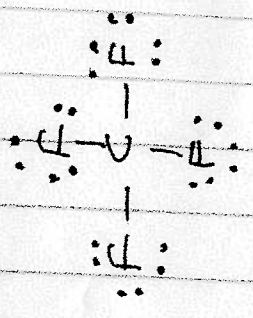
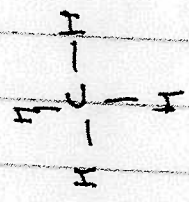
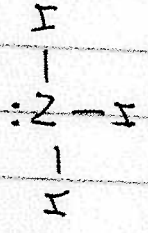
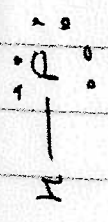
24

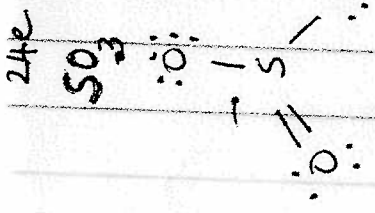
8

0

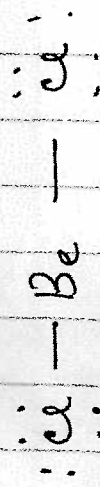
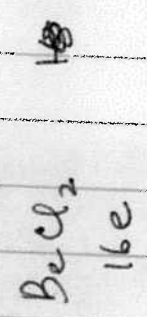
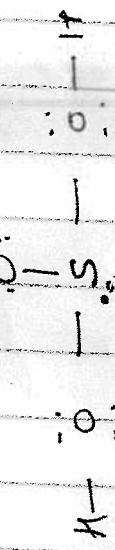
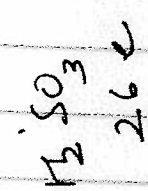
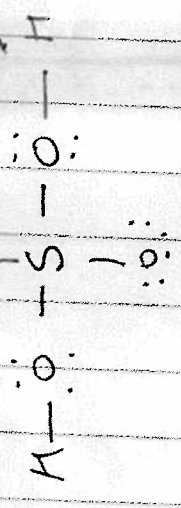
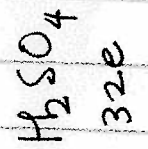
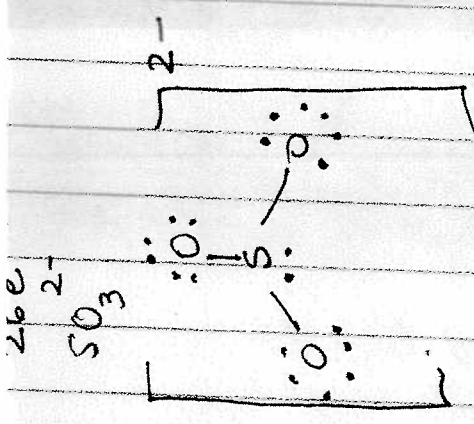
24e

Lewis st

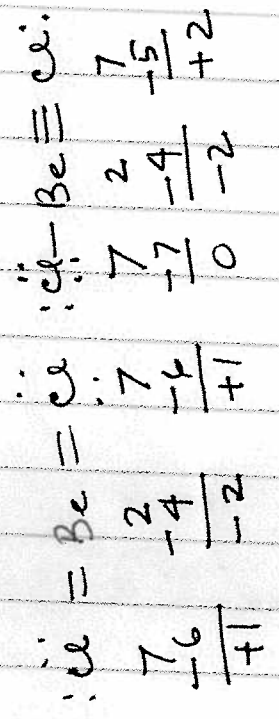
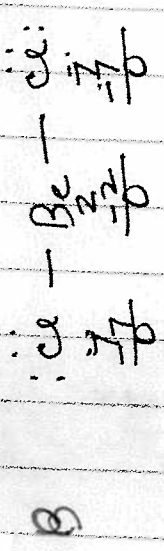
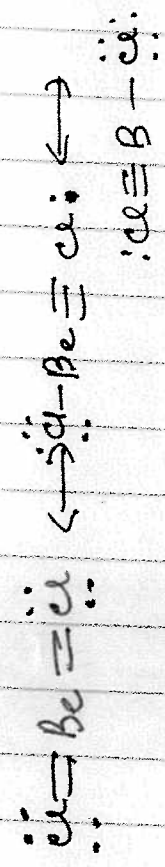




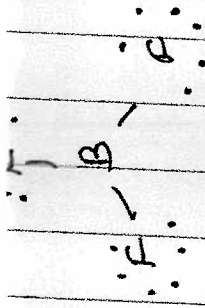
resonance structure  
smaller bond length



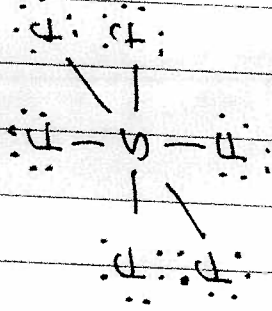
exception to octet rule



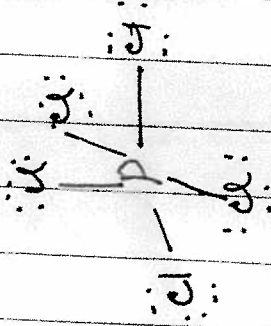
$$\frac{BF_3}{7 \times 3 + 3 = 24e}$$



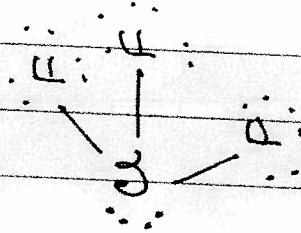
$$\frac{SF_6}{6 + 7 \times 6 = 48e}$$



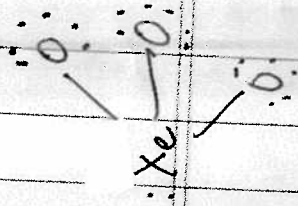
$$\frac{PCl_3}{5 + 5 \times 7 = 40e}$$



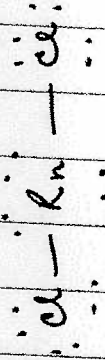
$$\frac{CF_3}{28e}$$



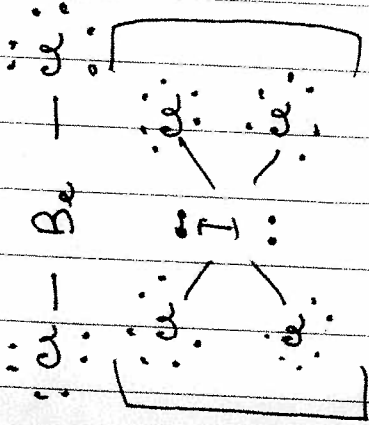
$$\frac{XeO_3}{8 + 18 = 26e}$$



$$\frac{RnCl_2}{8 + 14 = 22e}$$



$$\frac{BeCl_2}{2 + 7 \times 2 = 16e}$$



$$\frac{I_2Cl_4}{35 + 1 = 36e}$$

